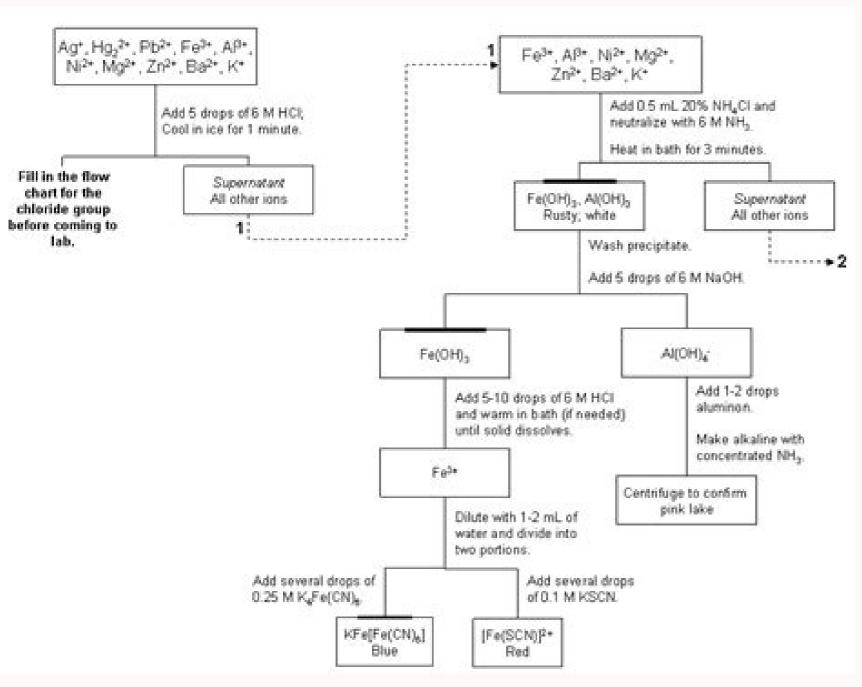
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Alternative approach for example 1

$$\begin{array}{ll} C(s) + O_2 \ (g) & \rightarrow & CO_2(g) \\ H_2(g) + \frac{1}{2}O_2(g) \rightarrow & H_2O(g) \\ C_3H_6(g) + 4\frac{1}{2}O_2(g) \rightarrow & 3H_2O(g) + 3CO_2(g)\Delta H^o_{298} = -2058.5 \ kJ \ mol^{-1} \\ \end{array}$$

$$3C(s) + 3H_2(g) \rightarrow C_3H_6(g)$$
 $\Delta H_f = ?$

Re-write the equations so that the reactants and products are on the same side of the "arrow" as the equation you are interested in. Multiply each equation so that there are the same number of moles of each constituent also.

$$3C(s) + 3O_{o}(a) \rightarrow 3CO_{o}(a)$$

 $\Delta Hc = 3 \times -394 \text{ kJ}$

 $3H_{2}(g) + 1\frac{1}{2}O_{2}(g) \rightarrow 3H_{2}O(g)$

 $\Delta Hc = 3 \times -286 \text{ kJ}$

 $3H_2O(g) + 3CO_2(g) \rightarrow C_3H_6(g) + 4\frac{1}{2}O_2(g)$

 $\Delta Hc = +2058.5 \, kJ$

Equation has been reversed; (enthalpy now has opposite sign)



ANALYTICAL CHEMISTRY.

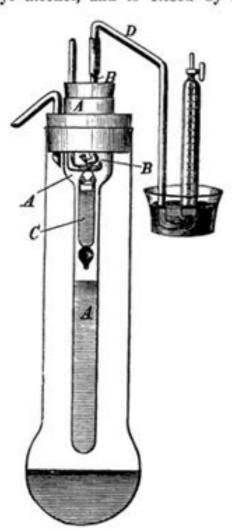
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phate" assisted the fermentations, and a very considerable loss of ammonia occurred. Lastly, experiment showed the importance of keeping the air from the dung, and that consequently less loss of ammonia occurs in the case of the dung being allowed to accumulate in the stalls than when it is regularly removed to a separate dungheap.

J. W. L.

Analytical Chemistry.

Apparatus for the Estimation of Nitrogen in Ammonium Salts. By W. Hentschel (Ber., 23, 2402—2403).—The method is based on the decomposition of ammonium salts by sodium hypobromite. The apparatus consists of an outer vessel similar to those employed in V. Meyer's vapour-density apparatus; it contains a small quantity of methyl alcohol, and is closed by a rubber stopper,



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